

STRUCTURE OF ATOM (ix)

1. How do electrons discovered?
2. What were canal rays?
3. What was Thomson's model of an atom and what was their main limitation?
4. On the basis of Rutherford's model of an atom, which sub atomic particle was discovered in the nucleus of an atom? How they conclude this?
5. Who discovered Neutrons? Draw a sketch of Bohr's model of an atom with three shells?
6. What are main rules to be followed for writing the number of electrons in different energy levels?
7. What do you understand by valency? Find the valency of Mg, N, O and P by their schematic atomic structure.
8. Write the relationship among no of electrons, protons and neutron and atomic no and atomic mass. Explain with the help of example.
9. What are isotopes and isobars? Explain.
10. In bromine atom available in the form of say, two isotopes $^{79}_{35}\text{Br}$ (49.7%) and $^{81}_{35}\text{Br}$ (50.3%) calculate the average atomic mass of Br.
11. The average atomic mass of a sample of an element x is 16.24 what are the percentage of isotopes $^{16}_8\text{x}$ and $^{18}_8\text{x}$ in the sample.