

PRACTICE TEST PAPER: CLASS-X

METALS AND NON-METALS

- Q1. Why do non-metals not conduct electricity? Also, name one of the non-metal that conducts electricity.
- Q2. Write balanced chemical equations for the following :
- Dilute sulphuric acid reacts with aluminium powder.
 - Dilute hydrochloric acid reacts with sodium carbonate.
 - Carbon dioxide is passed through lime water.
- Q3. Fill in the blanks:
- SO_2 is the product of of sulphide ores.
 - ZnO can be reduced by
 - forms greenish layer due to basic copper carbonate.
- Q4. a) Show the formation of NaCl from sodium and chloride atoms by the transfer of electrons.
- b) Why sodium chloride has a high melting point?
- Q5. What are amphoteric oxides? Choose the amphoteric oxides from Na_2O , ZnO , Al_2O_3 , CO_2 , H_2O .
- Q6. Give reason : Metals can be given different shapes according to our needs.
- Q7. a) Name the solvent in which the ionic compounds are generally soluble.
- b) Why are aqueous solutions of ionic compounds able to conduct electricity?
- Q8. From amongst the metals sodium, calcium, aluminium, copper and magnesium, name the metal
- Which reacts with metal only on boiling.
 - Another which does not react even with steam.
- Q9. Give reason : Gold and Silver are used for the making of jewellery.
- Q10. An element A reacts with water to form a compound B which is used in white washing. The compound B on heating forms an oxide which on treatment with water gives back B. Identify A,B and C and give the reactions involved.