

PRACTICE TEST PAPER: CLASS-X

ACIDS, BASES AND SALTS.

1. Name three natural and three synthetic indicators and their reactions with acids and bases.
2. What are olfactory indicators? Write an example, with the reaction.
3. Write balanced chemical equations for the following:
 - a. Sodium hydroxide + Zinc
 - b. Sodium Carbonate + Hydrochloric Acid
 - c. Sodium Hydrogen Carbonate + Hydrochloric Acid.
4. Name the chemical and write both the chemical equations-
On passing CO_2 gas, it turns milky. On passing excess carbon dioxide, white stone-like crystals are formed.
5. What do all acids and all bases have in common? Can it be concluded by any experiment? If yes, explain and give the experiment.
6. How is the concentration of hydronium ion (H_3O^+) affected when a solution of an acid is diluted?
7. Why does dry Hydrochloric Acid gas not change the color of the dry litmus paper?
8. While diluting an acid, why is it recommended that the acid should be added to the water and not water to the acid?
9. What are universal indicators? How does it conclude that given substance is acid or base?
10. How does pH change one of the causes of tooth decay?
11. What effect does concentration of H^+ (aq.) ions have on the nature of a solution?
12. What is brine? Explain chlor-alkali process.
13. Write balanced chemical equation for the making of baking soda.
14. What happens when baking soda is heated? What happens when the obtained substance is recrystallized? Give equations.
15. What happens when Plaster of Paris is mixed with water? Give chemical equation.